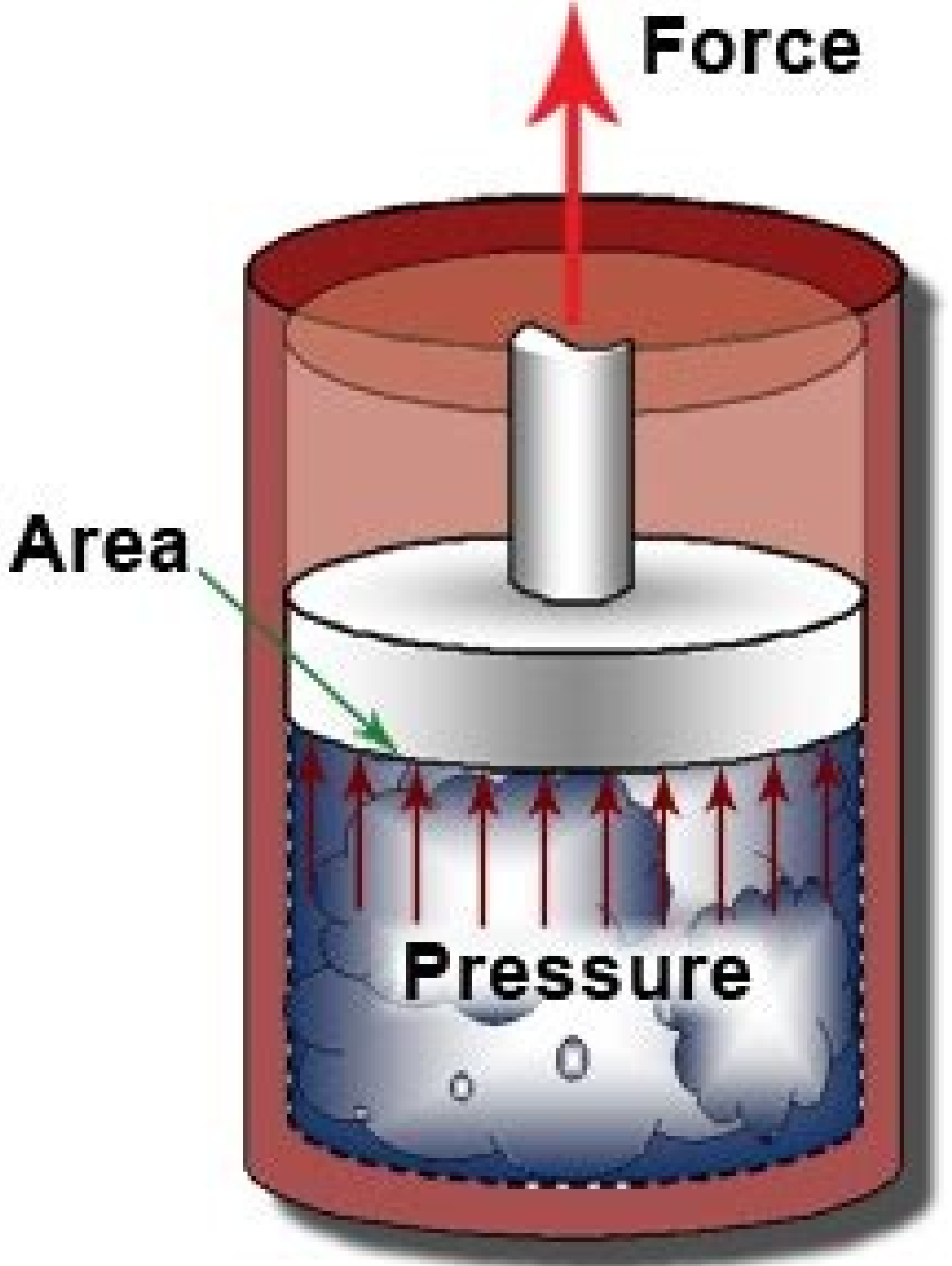


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$$w = \frac{1}{6} g$$

If $\frac{x}{y} = \frac{4}{5}$, $\frac{y}{z} = \frac{3}{10}$ and $\frac{z}{w} = \frac{6}{7}$, what is the value of $\frac{x+y+w}{z}$?

$5x = 4y$ $10y = 3z$ $7z = 6w$

$x = \frac{4}{5}y \Rightarrow y = \frac{3}{10}z$

$\Rightarrow x = \frac{4}{5} \times \frac{3}{10}z = \frac{12}{50}z$

$\frac{x+y+w}{z} = \frac{\frac{12}{50}z + \frac{3}{10}z + \frac{1}{6}z}{z} = \frac{12}{50} + \frac{3}{10} + \frac{1}{6} = \frac{128}{75}$

4 moles of nitrogen gas are confined to a 6.0 L vessel at 177 °AC and 12.0 atm. Paul Taylor, Getty Images A balloon contains 4 moles of an ideal gas with a volume of 5.0 L. What is the final volume? A mixture of helium and neon gases is held in a container at 1.2 atmospheres. What is the average velocity of nitrogen gas molecules at STP?Molar mass of nitrogen = 14 g/mol A 60.0 L tank of chlorine gas at 27 °AC and 125 atm springs a leak. What is the pressure of the gas? If the vessel is allowed to expand isothermally to 36.0 L, what would be the final pressure? If the mixture contains twice as many helium atoms as neon atoms, what is the partial pressure of helium? Helium gas has a molar mass of 2 g/mol. The temperature of a sample of an ideal gas in a sealed 5.0 L container is raised from 27 °AC to 77 °AC. At low temperatures, real gases behave like ideal gases. This collection of ten chemistry test questions deals with the concepts introduced with the ideal gas laws.Useful information:At STP: pressure = 1 atm = 700 mm Hg, temperature = 0 °AC = 273 KAt STP: 1 mole of gas occupies 22.4 LR = ideal gas constant = 0.0821 LA·atm/mol·AK = 8.3145 J/mol·AKAnswers appear at the end of the test. A 0.614 mole sample of ideal gas at 12 °AC occupies a volume of 4.3 L. Thank you very much for your cooperation. When the leak was discovered, the pressure was reduced to 50 atm. If the initial pressure of the gas was 3.0 atm, what is the final pressure? In order to continue enjoying our site, we ask that you confirm your identity as a human. A 9.0 L volume of chlorine gas is heated from 27 °AC to 127 °AC at constant pressure. The ideal gas law is an important concept in chemistry. What is the density (in g/L) of a gas with a molar mass of 60 g/mol at 0.75 atm and 27 °AC? How many moles of chlorine gas escaped? If an additional 8 moles of the gas is added at constant pressure and temperature, what will be the final volume of the Oxygen Gas has a molar mass of 32 g/mol.Would Cow much faster or slower have oxygen effusing from a small opening than helium? It can be used to predict the behavior of real gases in different situations of low temperatures or high pressures. pressures.

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